

EQUIVALENT MASS

Definition

The mass of a substance especially in grams that combines with or is chemically equivalent to **eight grams of oxygen** or **one gram of hydrogen**

$$\text{Equivalent Mass} = \frac{\text{Molar Mass}}{n\text{-Factor}}$$

EQUIVALENT MASS

Element

$$E = \frac{\text{Atomic mass of element}}{\text{Valence of element}}$$

Ion

$$E = \frac{\text{Formula mass of ion}}{\text{Charge on ion}}$$

Salt

$$E = \frac{\text{Formula mass of salt}}{\text{Total positive or negative charge on cationic or anionic part}}$$

Acid

$$E = \frac{M_{\text{acid}}}{\text{Basicity}}$$

Base

$$E = \frac{M_{\text{Base}}}{\text{Acidity}}$$

Acid Salt

$$E = \frac{M}{\text{Replaceable 'H' left in the salt}}$$

Redox Change

$$E = \frac{M_{\text{oxidant or reductant}}}{\text{Number of electrons lost or gained by one molecule of oxidant or reductant}}$$

